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'14-'15 Mr. Murdoch

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Solving for X, we get

+6, so the oxidation

number for sulfur in

the polyatomic ion (SO

4-2) is +6, and so

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sulfur will be +6 in
 ZnSO_4 . *Check the
Periodic Table to see if
sulfur has a +6 charge
listed; it does!

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Chemistry '14 Mr ...

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) A) 12 moles B) 2

moles C) 3 moles D)6

moles 41. Given the

balanced equation

representing a

reaction: $2\text{Fe} + 3\text{Cu}^{2+}$

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© $2\text{Fe}^{3+} + 3\text{Cu}$ When the iron atoms lose six moles of electrons, how many moles of electrons are gained by the copper ions? A) 1 mole B) 6 moles C) 3 moles D) 9 moles
42. Given the balanced equation:

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) 68. Base your answer
to the following
question on the
information and
diagram below. The
apparatus shown in the

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diagram consists of two inert platinum electrodes immersed in water. A small amount of an electrolyte, H_2SO_4 , must be added to the water for the ...

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Answers

Oxidation: loss of electrons
Reduction: gain of electrons
Oxidizing Agent: substance reduced

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Reducing Agent:

substance oxidized

Anode: electrode

where oxidation occurs

(AN OX) Cathode:

electrode where

reduction occurs (RED

CAT) Salt Bridge:

allows the flow of ions

between the anode

and cathode solutions

to keep a balance of

charge Cations flow

toward the Cathode

Chemistry Regents

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A 12 moles B 2 moles

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Answer Questions

Electrochemistry is the chemistry of electrochemical reactions; which deal with the relationship between electrical energy and chemical reactions.

Electrochemical reactions involve transfer of electrons and are essentially REDOX reactions.

Displacement and REDOX Reactions

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Experiment 1:-

Displacement

Reactions among

Metals. Procedure

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moles C) 3 moles D) 6

moles 41. Given the

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42. Given the balanced equation:

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ELECTROCHEMISTRY 3

- All Free elements are

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zero • Group 1 is +1 •
Group 2 is +2 •
Fluorine is -1 • H is
usually +1, and if so,
will be first written in
the chemical formula. •
O is almost always -2,
exceptions: H_2O_2 and
 OF_2 • In compounds
using polyatomic ions,
the first and middle
elements all have +
oxidation states. • The
sum of oxidation states
in a ...

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