

# Chemquest 20 Ionic Bonding Advanced

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7. Write formulas for the following ionic compounds. Note that each includes a polyatomic ion. a) lithium phosphate b) ammonium oxide c) barium hydroxide  $\text{Li}_3\text{PO}_4$   $(\text{NH}_4)_2\text{O}$   $\text{Ba}(\text{OH})_2$  d) calcium cyanide e) sodium chlorate f) potassium sulfate  $\text{Ca}(\text{CN})_2$   $\text{NaClO}_3$   $\text{K}_2\text{SO}_4$  8. In question 3, you were asked the difference between calcium nitride and ...

## Information : Polyatomic Ions

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### Information: Polyatomic Ions

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Therefore, the charge on each X is 1. Here's why: Remember Answer: + that S has a negative 2 charge. There are two "X" ions to balance out the negative 2 charge of S.

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